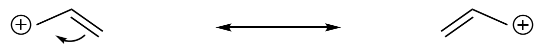
RECALL:

Conjugated Intermediates:

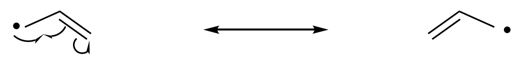
1) Cation:



Allyl Cation

Conjugation stabilizes the positive charge

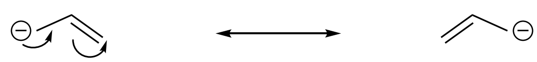
2) Radical:



Allyl Radical

Conjugation stabilizes the radical

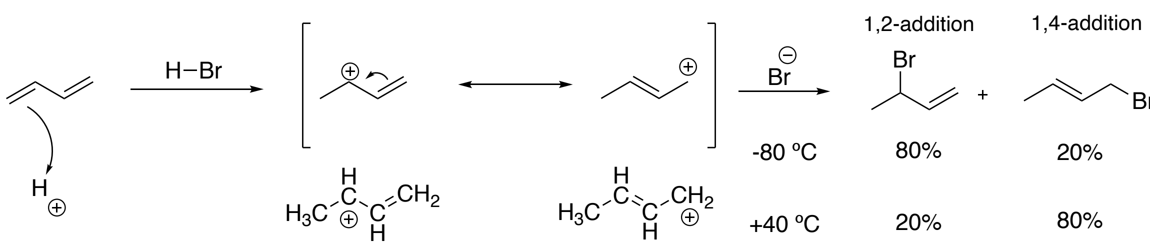
3) Anion:



Allyl Anion

Conjugation stabilizes the negative charge

Thermodynamic vs. Kinetic Products: Reaction of Dienes -



The two possible products are structural isomers.

The product 3-bromobutene is a **1,2-addition** product, whereas 1-bromo-2-butene is the product of a **1,4-addition** reaction. The numbers (**1,2-** or **1,4-**) indicate the position where the H and Br added to the 1,3-butadiene.

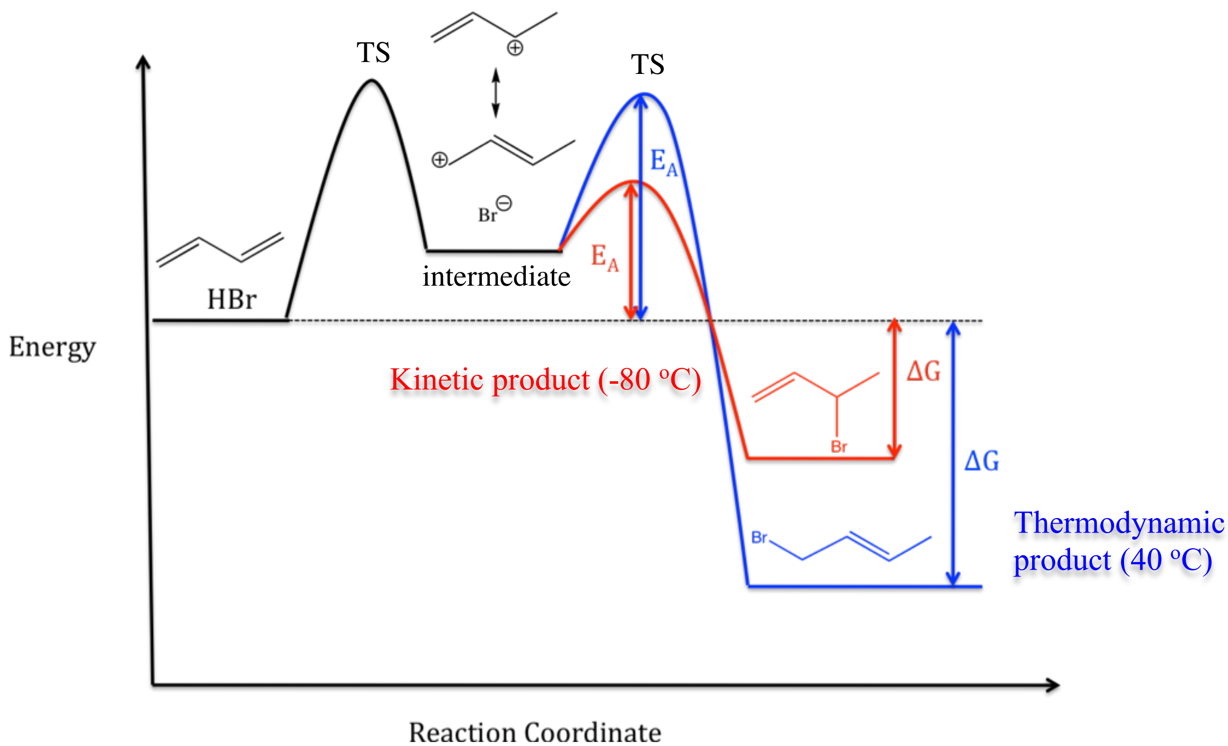
From the 2 resonance forms (connected by the double-headed arrow) of the allylic cation in the above figure, we see that the positive charge is shared between the C2 and C4.

Note: The allylic cation has two electrons delocalized across the three carbons. It was drawn to have two resonance forms, but it is a ***single entity***. Neither resonance form depicts the actual structure, but rather the molecule exists as a combination of the two resonance forms. The 'primary carbocation' is stabilized, as it is allylic.

Why does the temperature affect the ratio obtained?

- 3-bromo-1-butene has a higher yield at the lower temperature because it is formed faster than 1-bromo-2-butene due to its lower E_a: **kinetic control**
- 3-bromo-1-butene has lower yield than 1-bromo-2-butene at higher temperature due to **thermodynamic control**. The addition of bromine to the allylic cation is reversible at high temperature. 3-Bromo-1-butene can be converted back to the allylic cation, and then form 1-bromo-2-butene which is the thermodynamically more favoured product as it is more stable. The thermodynamic product is determined by the equilibrium result controlled by ΔG.

Now let's look at the energy diagram of the HBr addition to butadiene reaction



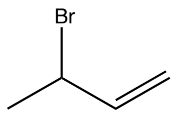
The activation energy barrier to form the 1,2-product, 3-bromo-1-butene is much *smaller* than the 1,4-product, 1-bromo-2-butene (so that it can be formed easier and **faster**. We call this **kinetically favored**).

However, the energy of 1-bromo-2-butene is lower than 3-bromo-1-butene, so that it is more **stable** than 3-bromo-1-butene (it is **thermodynamically favored**).

Kinetic control: governed by the activation energy (Ea)
Thermodynamic control: you get an equilibrium mixture. Governed by ΔG.

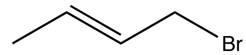
Why is 1-bromo-2-butene more stable?

- The bromine atom is bulky. It likes to stay away from the rest of the molecule to avoid steric clashes.
- Alkene carbons are somewhat electron deficient. More highly substituted alkenes are more stable due to donation of electron density by the substituents (C vs. H)



1,2-addition

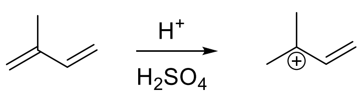
monosubstituted C=C



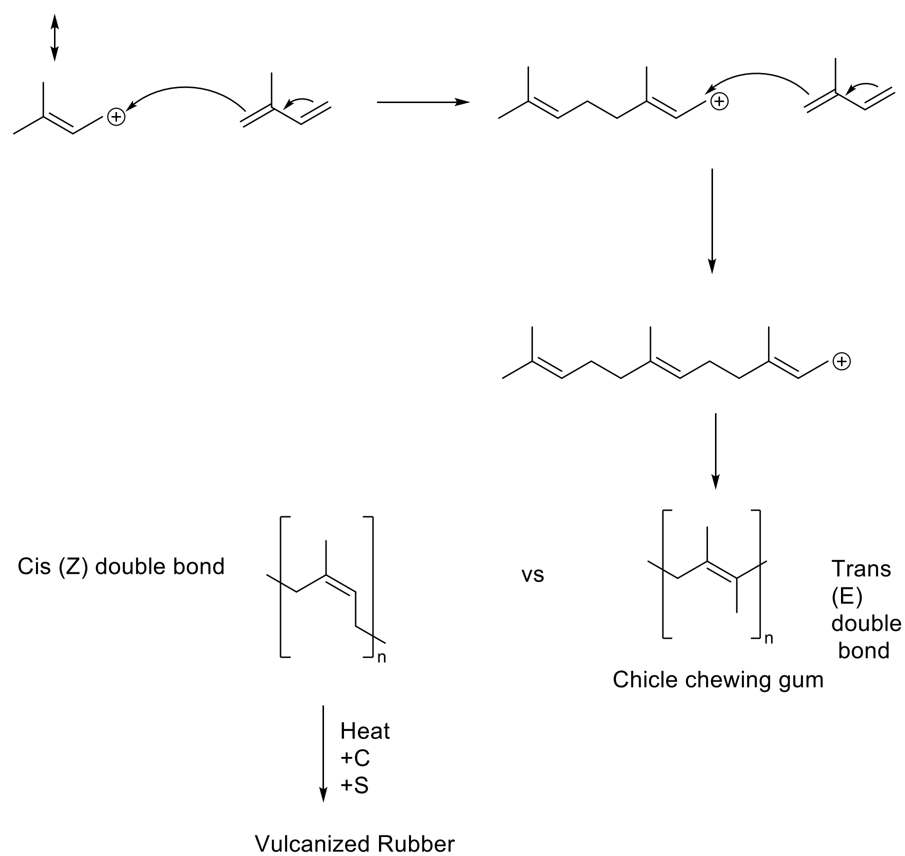
1,4-addition

disubstituted C=C

Polymerization: If no nucleophile is present in previous addition reaction - e.g. isoprene

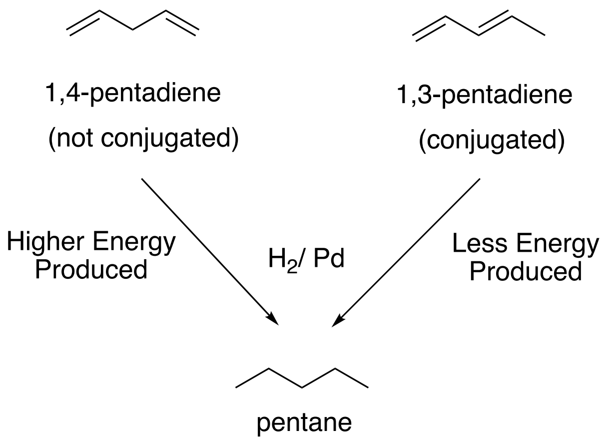


Protonation of dienes with strong acid can form a conjugated carbocation (has a resonance form)



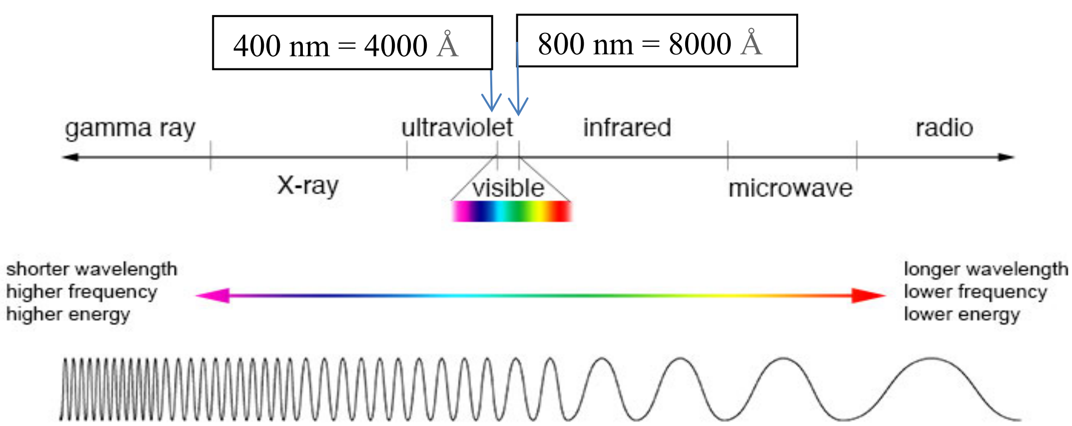
Rubber tree (*Hivea brasiliensis*) produces approximately 2000-4000 lbs/acre of rubber.

Example of conjugated and not conjugated system:



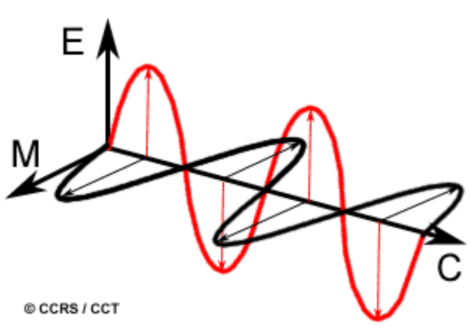
Electromagnetic Spectrum:

1nm = 10 angstrom



UV and visible light: conjugated double bond systems absorb UV light and some visible light

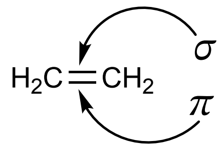
$$\mathbf{E = h\nu = \frac{hc}{\lambda}}$$



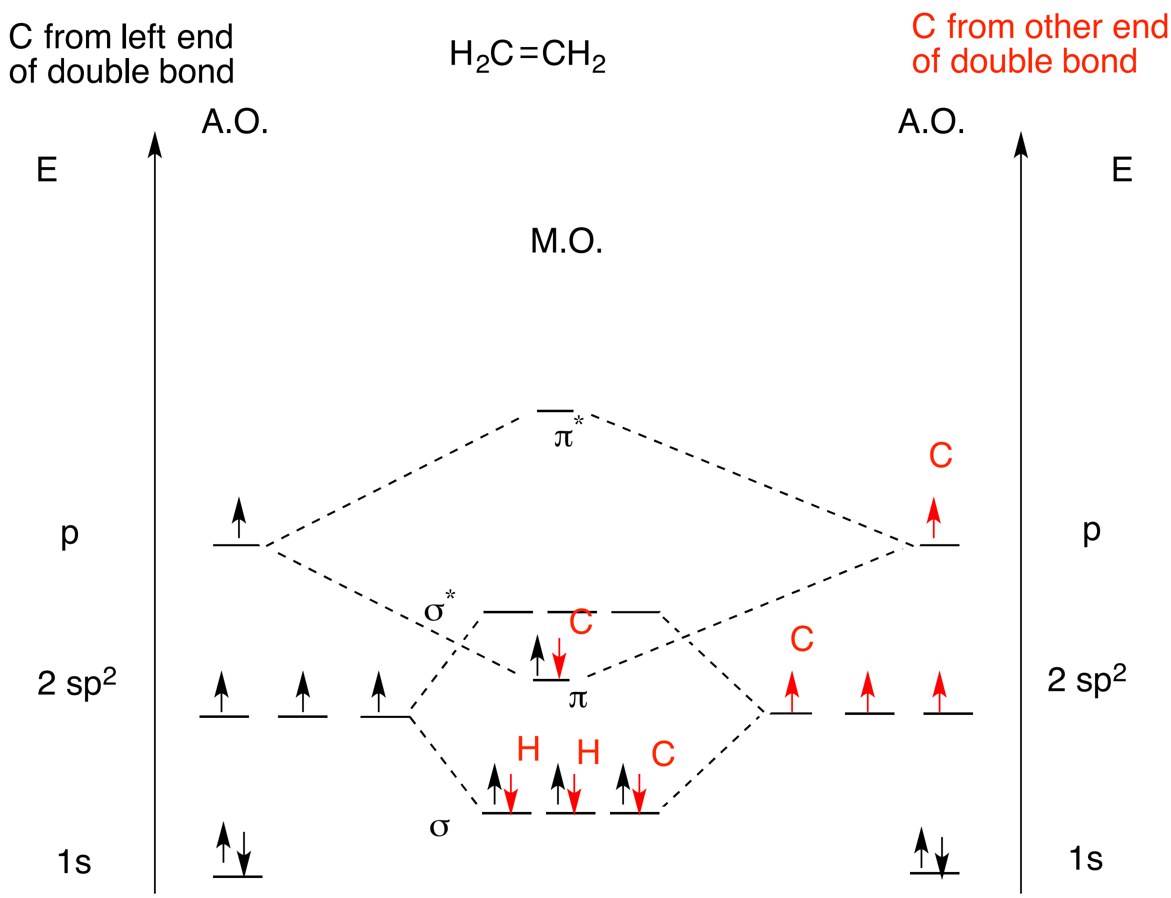
E = energy
H = Planck's Constant (6.6 x 10^{-34} Joules•sec
ν = frequency
λ = wavelength
c = speed of light (3.0 x 10^{10} cm/sec)

Molecular Orbitals:

Example 1: Ethylene

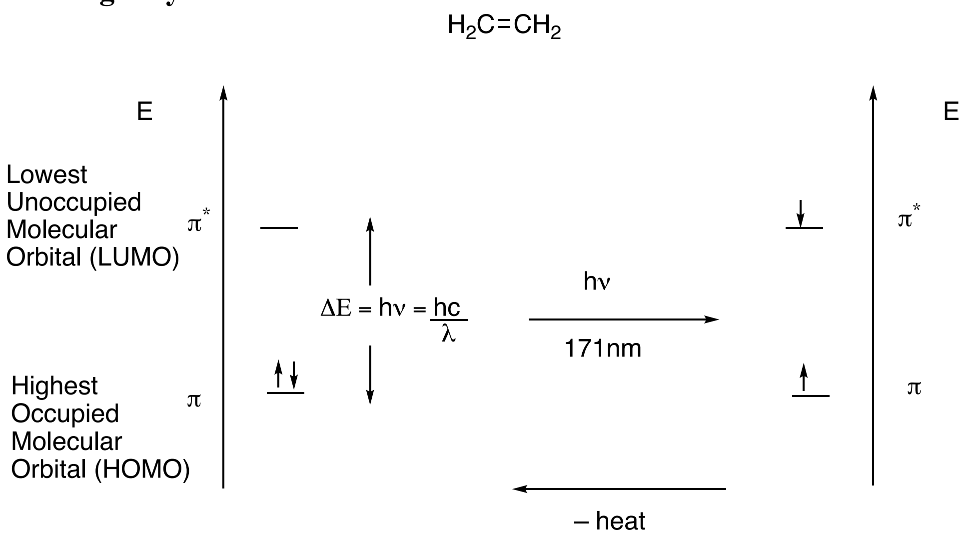


Looking both at sigma and pi orbitals:



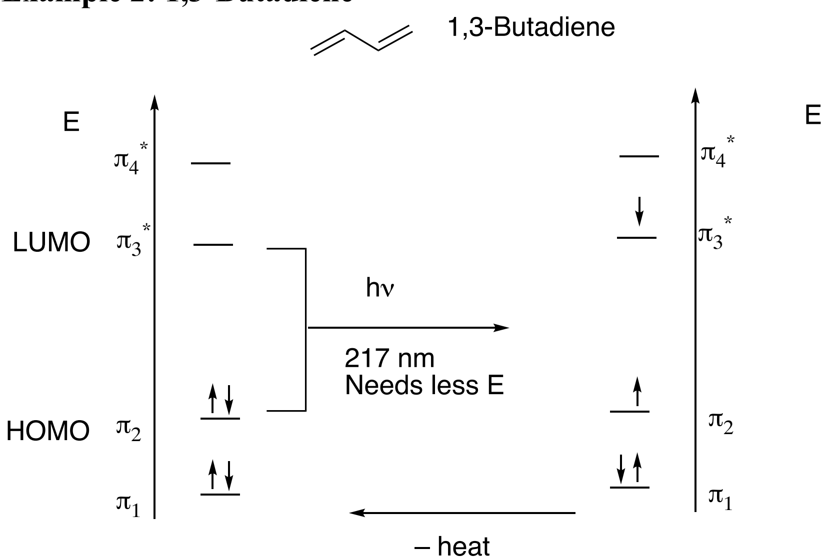
A.O. means atomic orbitals (s, sp^2, p)
M.O. means molecular orbitals (σ , π)

Looking only at the π orbitals:

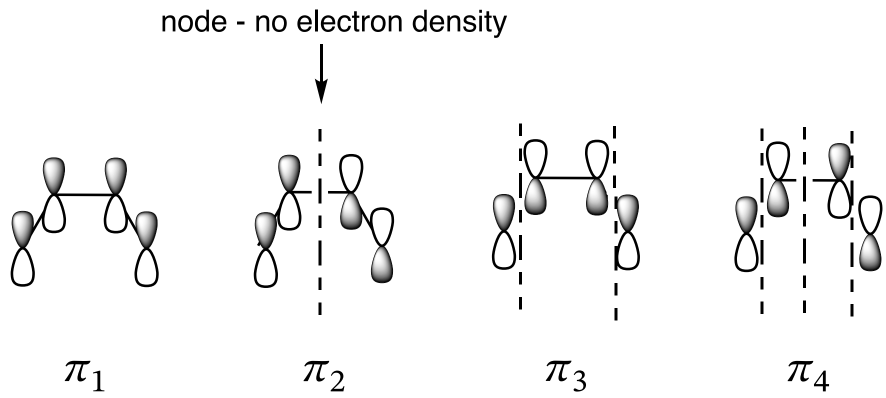


An electron can be excited from the HOMO to the LUMO using light of a precise wavelength dependent on the energy difference between the two orbitals (since the orbitals are quantized). The electron can go back to its original orbital and heat (or light) is produced in the process. When the electron is promoted to a higher energy state (excited to a higher energy molecular orbital), it attains a **singlet state**. The electron can go back to its original orbital and heat (or light) is produced in the process.

Example 2: 1,3-Butadiene



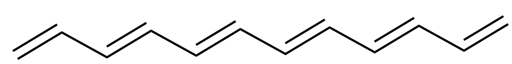
Node: a point or plane of zero electron density in an orbital
HOMO: Highest Occupied Molecular Orbital
LUMO: Lowest Unoccupied Molecular Orbital

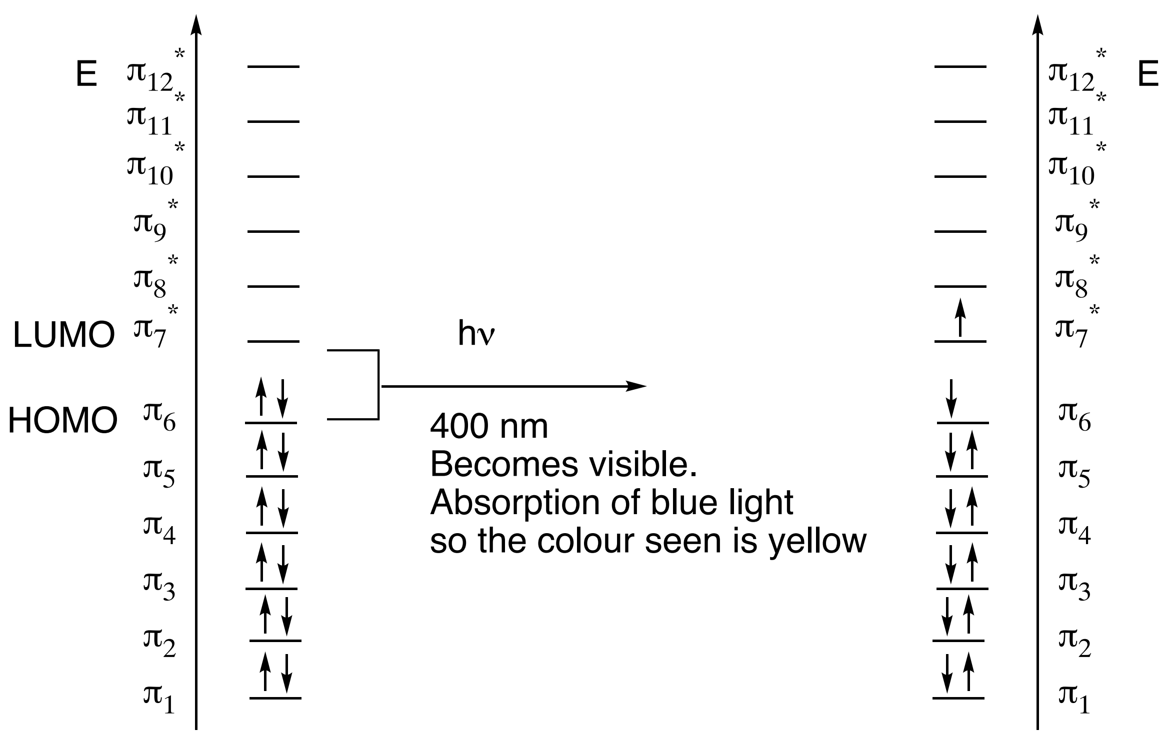


As the number of double bonds in the compound increases, decreasing the HUMO-LUMO gap, the energy of the light needed to get its excited state is lower.

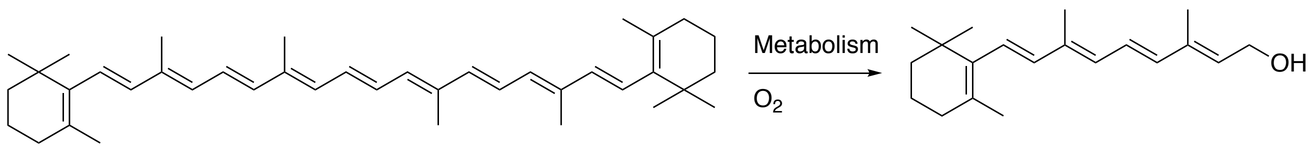
Once the absorption of light leaves the UV range and into the visible range, the transition becomes visible and the color of the compound can be seen.

Example 3: 1,3,5,7,9,11-dodecahexaene



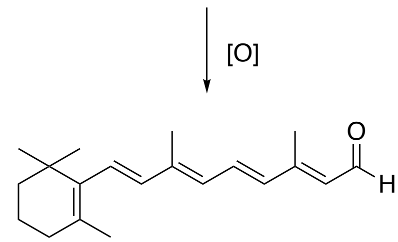


Conjugated molecules in vision:



β-Carotene

Vitamin A

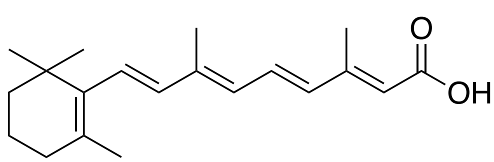


Retinal

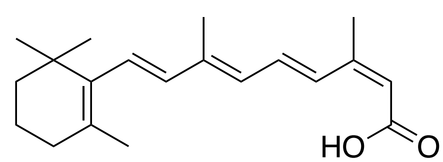
- Abundant in nature 10^9 metric tons
- Photoprotective agent. Orange pigment in plants.
- **Xanthophylls – oxidized carotenes**

Retinal is combined with the protein opsin in the eye to make rhodopsin, which is a key protein in the mechanism of sight.

Human vision covers the range of 400 nm (4000 Å) to 800 nm (8000 Å), anything outside of these wavelengths is invisible to the naked eye.
The further oxidized form of retinal is retinoic acid. Changing the double bond bearing the carboxylic acid from trans to cis gives the drug Accutane, used to treat acne. (Can cause birth defects = teratogen)



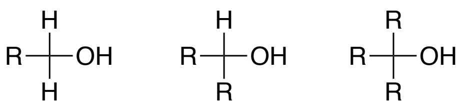
Retinoic acid



Accutane

Alcohol and Ether Nomenclature

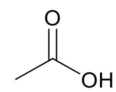
Alcohol	Ether
R-OH	R-O-R

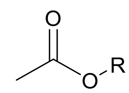


Primary Secondary Tertiary

1° 2° 3°

R ≠ 





Carbonyl **Acid** **Ester**

Alcohols are classified as primary (1°), secondary (2°), or tertiary (3°), depending on the number of organic groups bonded to the hydroxyl bearing carbon.

Note: –OH is called hydroxyl, hydroxy or alcohol

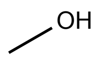
Naming:

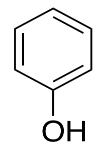
1. Find the longest chain, with the maximum number of OH groups.
2. Number in such a way to give the **first OH** the lowest number
3. Drop the "e" of the alkane name, add "ol"

Note: the alcohol (-OH) takes priority over ethers, multiple bonds, and halogens

Examples:

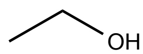
H_3C-OH
methanol
-toxic
wood alcohol





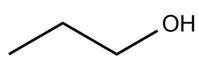
Hydroxybenzene
Phenol
C_6H_6O

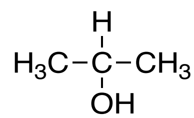
CH_3CH_2OH
ethanol
grain alcohol



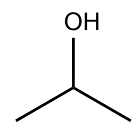
$CH_3CH_2CH_2OH$

1-propanol
n-propanol
propan-1-ol
propyl alcohol

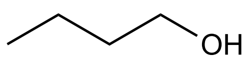




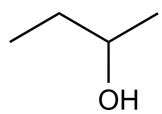
2-propanol
propan-2-ol
Isopropyl alcohol
isopropanol



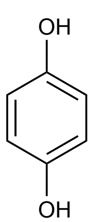
2° alcohol



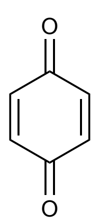
1-butanol
n-butanol



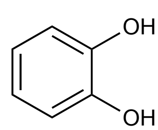
2-butanol



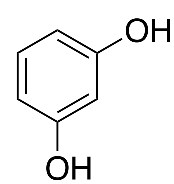
hydroxyquinone



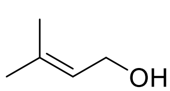
quinone



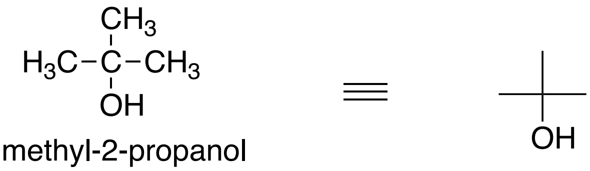
catechol



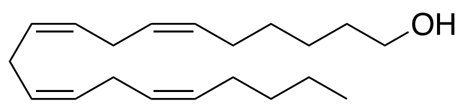
Resorcinol



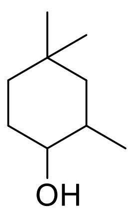
3-methyl-2-buten-1-ol



2-methyl-2-propanol
tertiary butyl alcohol
t-BuOH
tert-butyl alcohol

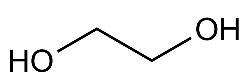


Eicosa-6Z,9Z,12Z,15Z-tetraen-1-ol

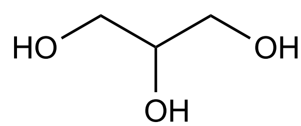


2,4,4-trimethylcyclohexanol

Polyols



ethan-1,2-diol
ethylene glycol



1,2,3-propantriol
glycerol